



## COURSE DESCRIPTION CARD - SYLLABUS

Course name

Elementy preparatyki nieorganicznej (Elements of inorganic preparation)

### Course

Field of study

Year/Semester

Technologia chemiczna (Chemical Technology)

II/3

Area of study (specialization)

Profile of study

general academic

Level of study

Course offered in

First-cycle studies

Polish

Form of study

Requirements

full-time

elective

### Number of hours

Lecture

Laboratory classes

Other (e.g. online)

0

30

0

Tutorials

Projects/seminars

0

0

### Number of credit points

3

### Lecturers

Responsible for the course/lecturer:

dr hab. eng. Bogdan Wyrwas

Responsible for the course/lecturer:

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### Prerequisites

Knowledge:

Student: has knowledge resulting from passing the subject in general and inorganic chemistry during the 1st and 2nd semester, and in particular:

W1. Student has established theoretical knowledge in the field of inorganic and general chemistry, and in particular, is able to describe the structure of matter at the nuclear, atomic and molecular level; can identify properties of elements and their compounds being able, at the same time, to explain these properties in relation to the position of elements in the periodic table



W2. Student knows the EHS of chemistry laboratory, especially the rule of keeping the workplace clean; knows First Aid basic principles and is able to apply them in the event of unfortunate accidents and incidents

W3. Student can list and characterize basic laboratory techniques

W4. Student is able to plan and carry out simple chemical experiment and knows how to analyze, develop and describe its results

Skills:

Student: has skills resulting from passing the subject in general and inorganic chemistry during the 1st and 2nd semester, and in particular:

U1. Student is skilled in chemical calculations: is able to use the periodic table: can write molecular and structural formulas of chemical compounds as well as can write and balance chemical reactions of any type involving inorganic compounds

U2. Student is able to analyze and solve typical chemical problems based on knowledge from various sources, including knowledge gained by his/her own; knows how to compare knowledge from different sources

U3. Student can organize his/her own work in the chemistry laboratory; can correctly apply different laboratory techniques; can make proper use of laboratory equipment and correctly interpret obtained results

Social competences:

Student: has knowledge resulting from passing the subject in general and inorganic chemistry during the 1st and 2nd semester, and in particular:

K1. Student can see the dependence between his/her own safety and the safety of other people working in the laboratory in proceeding according to the regulations applied in the laboratory; develops the habit of keeping the workplace clean

K2. Student is aware of the adverse effects some commonly used inorganic compounds may have on environment; understands the need to make action required to minimize these harmful effects

### Course objective

Broadening of knowledge and improvement of practical skills related to laboratory work. Understanding the importance of being obedient to the EHS laboratory rules and any other restrictions; broadening of knowledge on the novel methods and techniques in use; team orientation for better work performance; transfer of knowledge within the practical application of chemical reactions and the related problem



solving techniques; getting familiar with the chemical transformation effects and getting practical knowledge on these effects

### Course-related learning outcomes

#### Knowledge

1. Student has established theoretical knowledge and laboratory practice in inorganic and general chemistry; can identify the properties of elements and their compounds, and is able to select proper methods and measures necessary for the practical application of chemical reactions and/or operations (K\_W03, K\_W08)
2. Student can identify the character of various chemical reactions and processes, which allows him/her to recognize specific problems that may occur during their application, especially when he/she is the person in charge when working in team; student is familiar with the laboratory EHS and obeys these rules unconditionally (K-W18)
3. Student can distinguish, characterize and explain the specific character of various techniques applied in laboratory; Student can plan complex chemical experiments and select all the necessary measures and techniques for them to be applied.; student knows how to apply the results (K\_W15)
4. Naming proper physicochemical properties of elements and their compounds, student can not only explain the simple - one-step -possibility for preparation of some compounds ( i.e. by conducting the reaction or chemical operation), but also the necessity of applying multi-stage procedures in other cases (K\_W03, K\_W08)

#### Skills

1. Student is skilled in writing and balancing chemical reactions of any type involving inorganic compounds and their thermodynamic conditions; student can use his/her in practice while conducting reactions and complex experiments involving inorganic reagents (K\_U01, K\_U18)
2. Student is able to analyze and solve typical chemical problems based on knowledge from various sources, including knowledge gained by his/her own; knows how to compare knowledge from different sources; has team-oriented attitude when complex chemical issues are developed and planned (K\_U01, K\_U02, K\_U16)
3. Student can organize his/her own work in the chemistry laboratory; can correctly apply different laboratory techniques; can make proper use of laboratory equipment and correctly interpret obtained results (K\_U01, K\_U07, K\_U20)
4. Students implements and follows the EHS rules of chemistry laboratory (K\_U10, K\_U28)

#### Social competences

1. Student can see the dependence between his/her own safety and the safety of other people working in the laboratory in proceeding according to the regulations applied in the laboratory; develops the habit of keeping the workplace clean (K\_K03, K\_K04)



2. Student is aware of specific risks and responsibilities he/she must take for implementation of entrusted tasks while working in a team (K\_K01, K\_K03, K\_K05)

### Methods for verifying learning outcomes and assessment criteria

Learning outcomes presented above are verified as follows:

The teacher regularly controls the theoretical preparation of students to perform planned exercises. The check is carried out by polling and / or in the form of written tests. The teacher observes and assesses the behavior of students in the laboratory, including the ability to organize laboratory work and their manual skills during the performance of the exercises planned. Written reports on performed exercises are subject to evaluation. The final grade is the result of the three components listed above

### Programme content

1. Obtaining salt by dissolving metals in acids
2. Reconstitution of amphoteric metals in concentrated alkali
3. Preparative use of exchange reaction to obtain sparingly soluble compounds
4. Precipitation of a sparingly soluble pellet as an intermediate step provided that the substrate and the product are well soluble in water
5. Other oxidants than hydrogen ion and anions derived from strong acids in inorganic preparation
6. Preparation of double salts and other compound mixtures
7. Preparation of sparingly soluble metal oxides and hydroxides
8. Preparative reactions with the participation of gaseous reactants

### Teaching methods

Classes are practical, they consist in the students themselves doing exercises included in the course plan. Each student is required to first submit their own concept of the assigned task, which is approved by the teacher. With the teacher's consent, it is allowed to carry out the additional tasks proposed by the students themselves. The teacher personally shows and explains how to perform the activities and operations that students meet for the first time. The teacher constantly controls the student's behavior in the laboratory and the way of performing his work themselves. He immediately notices and corrects irregularities

### Bibliography

Basic

1. J. Supniewski, Preparatyka nieorganiczna, PWN, Warszawa 1958
2. J. Gałęcki, Preparatyka nieorganiczna, WNT, Warszawa 1964
3. A. Bielański, Podstawy chemii nieorganicznej, t.1-3, PWN, Warszawa 2005



4. L. Jones, P. Atkins, Chemia ogólna. Częsteczki, materia, reakcje, tom 1 i 2, PWN, Warszawa 2009

#### Additional

1. A. Ciszewski, M. Baraniak, Aktywność chemiczna i elektrochemiczna pierwiastków w środowisku wody, Wydawnictwo PP, Poznań 2006

2. L. Kolditz, Chemia nieorganiczna, PWN, Warszawa 1994

3. F. Domka, J. Jasiczak, Analiza jakościowa, Wydawnictwo AE, Poznań 2004

4. K. M. Pazdro, Zbiór zadań z chemii, Oficyna Edukacyjna 2007

5. M.J. Sienko, R.A. Plane, Chemia. Podstawy i zastosowania, WNT, Warszawa 2002

#### Breakdown of average student's workload

	Hours	ECTS
Total workload	75	3,0
Classes requiring direct contact with the teacher	40	1,6
Student's own work (literature studies, preparation for laboratory classes, preparation for tests, preparation of laboratory reports) <sup>1</sup>	35	1,4

<sup>1</sup> delete or add other activities as appropriate